**Must Know List For TBL (Chemical Kinetics – Part I)**

1. Express the rate of a chemical reaction in terms of changes in concentrations of reactants and products with time.

- Study Example 16-1

2. Describe the experimental factors that affect the rate of chemical reactions, especially:

(i) Nature of Reactants

(ii) Concentration of Reactants

- Describe and use the rate-law expression for a reaction (the relationship between concentration and rate). Take note of definitions and of the points about the specific rate constant, *k*.

- Apply the method of initial rates to find the rate-law expression for a reaction.

- Study Examples 16-2, 16-3 & 16-4.