**COSTAATT**

**CHEM092 – Introduction to Concepts in Chemistry II**

**Lesson 10 – Worksheet**

1. For each of the following electrolytes: (i) write the cathode equation; (ii) write the anode equation; (iii) say what has been oxidised and what has been reduced.

(a) molten lead(II) bromide using carbon electrodes

(b) sodium chloride solution using carbon electrodes

(c) copper(II) sulphate solution using carbon electrodes

(d) copper(II) sulphate solution using copper electrodes

2. If molten CuCl2 is electrolyzed using inert electrodes, what substances will be produced at the anode and the cathode?

3. An aqueous 1M solution of NiSO4 was electrolyzed using inert electrodes. What

substance was produced at each electrode?

4. (a) What is the function of electroplating?

(b) Give the equation that occurs at the anode and the cathode during chromium plating.

(c) Draw a diagram to show a piece of steel being chrome plated.