**COSTAATT**

**CHEM 111**

**Final Revision Worksheet**

**Lecturer: Ms. Romona Olton**

**Student Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**SECTION A – Multiple Choice**

**1.** What is the mass of calcium in 100g of calcium carbonate, CaCO3?

(Relative Atomic masses: Ca = 40; C = 12; O = 16)

(A) 40g (B) 20g (C) 12g (D) 100g

**2.** Calculate the relative formula mass of iron(III) oxide, Fe2O3.

(Relative Atomic masses: Fe = 56; O = 16)

(A) 128g (B) 104g (C) 72g (D) 160g

**3.** Calculate the percentage of aluminium in aluminium oxide, Al2O3.

(Relative Atomic masses: Al = 27; O = 16)

(A) 62.8% (B) 26.5% (C) 52.9% (D) 42.9%

**4.** Metal ions always have a charge that is:

(A) Negative (B) Neutral (C) Positive (D) Variable

**5.** A big number 3 in front of a formula means:

(A) the same as a small number 3

(B) you multiply the formula by 3

(C) you must add 3

(D) you only multiply the atom next to the number 3 by 3

**6.** The formula CaCO3 means:

(A) three atoms of everything

(B) three carbon atoms and three oxygen atoms

(C) one calcium atom, one carbon atom and three oxygen atoms

(D) a bracket is missing

**7.** The equation below represents the reaction of sodium metal with water.

Magnesium + Water  Magnesium hydroxide + Hydrogen

Which of the following chemical equations represents a complete balanced chemical

equation for the given word equation?

(A) 2Mg(s) + 2H2O(l)  Mg2(OH)2(aq) + H2(g)

(B) Mg(s) + 2H2O(l)   Mg(OH)2(aq) + 2H2(g)

(C) 2Mg(s) + 2H2O(l)  2MgOH(aq) + 2H2(g)

(D) Mg(s) + H2O(l)  MgO(aq) + H2(g)

**8.** Ferric oxide reacts with aluminium to produce aluminium oxide and iron. The balanced

chemical equation for the given reaction is:

Fe2O3 + 2Al  Al2O3 + 2Fe

Which of the following substances is reduced in the given reaction?

(A) Al2O3 (B) Fe2O3 (C) Al (D) Fe

**9.** White silver chloride changes to grey in sunlight because of the formation of silver

metal along with the evolution of chlorine gas. Identify the chemical equation that

gives a correct representation of the process.

(A) 2AgCl  2Ag + Cl2 (B) AgCl Ag + Cl2

(C) 2Ag + Cl2  2AgCl (D) Ag + Cl2 2AgCl

|  |
| --- |
| **10.** What is the colour of litmus in alkali?(A) red (B) green(C) colourless (D) blue  |
| **11.** Which of the following is a property of all nitrates?(A) They are soluble(B) They are coloured(C) They are insoluble(D) They are hydrated |

**12.** When copper oxide is treated with hydrogen gas copper is produced along with water.

The balanced chemical equation for the given reaction is:

CuO + H2 Cu + H2O

Which substance is oxidized in the given reaction?

(A) H2 (B) Cu (C) H2O (D) CuO

**13.** What is it called when one compound breaks down into two or more compounds?

(A) displacement

(B) synthesis

(C) decomposition

(D) neutralization

**14.** What is the balanced chemical equation for the reaction of nitrogen with hydrogen to produce

ammonia gas?

(A) N2(g) + H2(g)  2NH3(g) (B) 2N2(g) + H2(g) 2NH3(g)

(C) N2(g) + 3H2(g)  2NH3(g) (D) 2N2(g) + H2(g)  2NH3(g)

**15.** In which reaction is sulphur dioxide acting as an oxidising agent?

(A) SO2 + 2H2O + Cl2 → H2SO4 + 2HCl

(B) SO2 + 2NaOH → Na2SO3 + H2O

(C) 2SO2 + O2 → 2SO3

(D) SO2 + 2H2S → 2H2O + 3S

**16.** Which of the following reactions represents a combination reaction?

(A) CaO(s) + H2O(g) Ca(OH)2(aq) (B) CaCO3(s) CaO(s) + CO2(g)

(C) Zn(s) + CuSO4(aq) ZnSO4(aq) + Cu(s) (D) 2FeSO4(s) Fe2O3(s) + SO2(g) + SO3(g)

**17.** Which of the following reactions represents a double displacement reaction?

 (A) BaCl2(aq) + Na2SO4(aq) BaSO4(s) + 2NaCl(aq)

(B) 2FeSO4(s) Fe2O3(s) + SO2(g) + SO3(g)

(C) 2Pb(NO3)2(g) 2PbO(s) + 4NO2(g) + O2(g)

(D) Zn(s) + 2AgNO3(aq) Zn(NO3)2(aq) + 2Ag(s)

**18.** What is the general formula for an alkane?

(A) CnH2n+2 (B) CnH2n (C) CnH2n+4 (D) CnH3n

**19.** What is the name of the straight chain alkane containing six carbon atoms?

(A) sextane (B) hexane

(C) pentane (D) heptane

**20.** What is the name of the organic compound shown below?



(A) pentanoic acid

(B) butanoic acid

(C) propanoic acid

(D) hexanoic acid

**21.** What is the functional group in the compound shown below?



(A) ethanol

(B) hydroxyl

(C) CnH2n+1OH

(D) homologous series

**22.** Which of the following numbers represents a strongly acidic pH?

(A) 1 (B) 5 (C) 11 (D) 14

|  |
| --- |
| **23.** Which of the following numbers represents an alkaline pH?(A) 1 (B) 7 (C) 5 (D) 11 |
| **24.** What is the most abundant gas in the atmosphere? |

(A) oxygen

(B) nitrogen

(C) carbon dioxide

(D) water vapour

**25.** Given the redox reaction:

    2I-(aq) + Br2(l) -> 2Br-(aq) + I2(s)

What occurs during this reaction?

(A) The I- ion is oxidized, and its oxidation number increases.

(B) The I- ion is oxidized, and its oxidation number decreases.

(C) The I- ion is reduced, and its oxidation number increases.

(D) The I- ion is reduced, and its oxidation number decreases.

 **(25 Marks)**

**SECTION B – Structured Questions**

**1.** The two isotopes of chlorine are shown below.



**(a)** What are isotopes?

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 **(1 Mark)**

**(b)** Complete the table below to show the number of subatomic particles in each

isotope of chlorine.

|  |  |  |
| --- | --- | --- |
|  |  |  |
| Number of Protons |  |  |
| Number of Neutrons |  |  |
| Number of Electrons |  |  |

 **(6 Marks)**

**(c)** Draw a shell diagram to show the electronic configuration of the isotope of chlorine which has a mass number of 35.

 **(3 Marks)**

**(d)** Write the electronic configuration for the isotope of chlorine drawn in part (c).

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 **(1 Mark)**

**(e)** When chlorine undergoes ionic bonding it forms the Cl**-** ion. Write the electronic configuration for the Cl**-** ion.

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 **(1 Mark)**

**2.** Athletes sometimes take drugs because the drugs improve their performance.

**(a)** One of these drugs is ephedrine. Ephedrine has the formula:

**C10H15NO**

Calculate the relative formula mass of ephedrine. Show clearly how you work out your answer.

Relative atomic masses: H = 1; C = 12; N = 14; O = 16.

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 **(2 Marks)**

 **(b)** Another drug is amphetamine which has the formula:

**C9H13N**

The relative molecular mass (*M*r) of amphetamine is 135.

Calculate the percentage by mass of nitrogen in amphetamine.

(Relative atomic mass: N = 14)

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 **(3 Marks)**

**(c)** In an experiment a chemist calculated that the maximum yield of amphetamine is 200g.The chemist did the experiment but only made 120g of amphetamine.

Calculate the percentage yield of amphetamine for this experiment.

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  **(3 Marks)**

**3.** Use oxidation numbers to determine if the following reactions are redox reactions.

If they are redox reactions, give the reactant which has been oxidised and the reactant which has been reduced:

**(a)** 2KI(aq) + Cl2(g) 2KCl(aq) + I2(aq)

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**(b)** 2Fe(s) + 3Cl2(g)  2FeCl3(s)

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**(c)** FeCl2(aq) + (NH4)2S(aq)  FeS(s) + 2NH4Cl(aq)

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**(d)** BaCl2(aq) + H2SO4(aq) BaSO4(s) + 2HCl(aq)

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 **(8 Marks)**

**4.** In each of the following reactions give the reducing and the oxidising agents:

**(a)** 2Fe2O3(s) + 3C(s) 3CO2(g) + 4Fe(s)

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**(b)** 2NO(g) + 2CO(g) N2(g) + 2CO2(g)

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**(c)** Zn(s) + 2HCl(aq) ZnCl2(aq) + H2(g)

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 **(6 Marks)**

**5.** **(a)** Draw the fully displayed formula and write the names of the following hydrocarbons:

(i) CH3CH2CH2OH (ii) CH3CH2CH2 CH2CH2CH2CH3

 (iii) CH3CH2CH2CH2CH2CH=CH2 (iv) CH3OH

 **(8 Marks)**

**(b)** Look carefully at the four compounds in part (a). Which two belong to the same homologous series?

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 **(1 Mark)**

**(c)** What is the functional group of the compound shown in 3(a)(iv)?

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 **(1 Mark)**

**(d)** What do you understand by the term ‘structural isomerism’?

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 **(2 Marks)**

**(e)** Draw and name a structural isomer of the compound shown in 3(a)(ii).

 **(3 Marks)**

**Total = \_\_\_\_\_\_/ 75 = \_\_\_\_\_\_\_\_\_%**