**CHEM132**

**TBL – Solutions – Part I**

**Must Know List**

1. Describe the factors that favor the dissolution process.

2. Describe the dissolution of solids in liquids, liquids in liquids and gases in liquids.

3. Describe the factors that affect rate of dissolution and saturation.

4. Describe how temperature and pressure affect solubility.

5. Express concentrations of solutions in molality units and as mole fractions (See Examples

14-1, 14-2 & 14-3).

6. Explain what is meant by the term colligative properties and list the four important colligative properties of solutions.

7. Carry out calculations involving the four colligative properties of solutions: lowering of vapor pressure (Raoult’s Law), boiling point elevation, freezing point depression, and osmotic pressure (See Examples 14-4, 14-5 & 14-6).